Updated: 23 March 2020

Print version

CEE 680: Water Chemistry

Lecture #30

Coordination Chemistry: case studies

(Stumm & Morgan, Chapt.6: pg.305-319)

Benjamin; Chapter 8.1-8.6

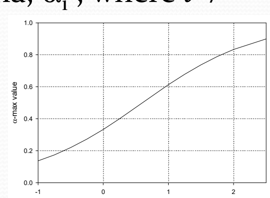
Summary of a few general rules

- At the intersection of sequential alphas, α_i and α_{i+1} :

$$Log[L] = pK_{i+1}$$

- At the peak of an intermediate alpha, α_i , where $i \neq 0$, *or the coordination number*

$$Log[L] = \frac{1}{2}(pK_i + pK_{i+1})$$



$LogK_i - LogK_{i+1}$

- This peak is also usually near intersection of the previous and following alphas (i.e., α_{i-1} and α_{i+1}), and its maximum height is estimated from:

Cadmium Complexes

- Bisulfide Ligand HS^-

Species	log K	Log Beta
CdL	Log K_1 = 10.17	Log β_1 = 10.17
CdL_2	Log K_2 = 6.36	Log β_2 = 16.53
CdL_3	Log K_3 = 2.18	Log β_3 = 18.71
CdL_4	Log K_4 = 2.19	Log β_4 = 20.90

- Cyanide Ligand CN^-

Species	log K	Log Beta
CdL	Log K_1 = 5.32	Log β_1 = 5.32
CdL_2	Log K_2 = 5.05	Log β_2 = 10.37
CdL_3	Log K_3 = 4.46	Log β_3 = 14.83
CdL_4	Log K_4 = 3.46	Log β_4 = 18.29

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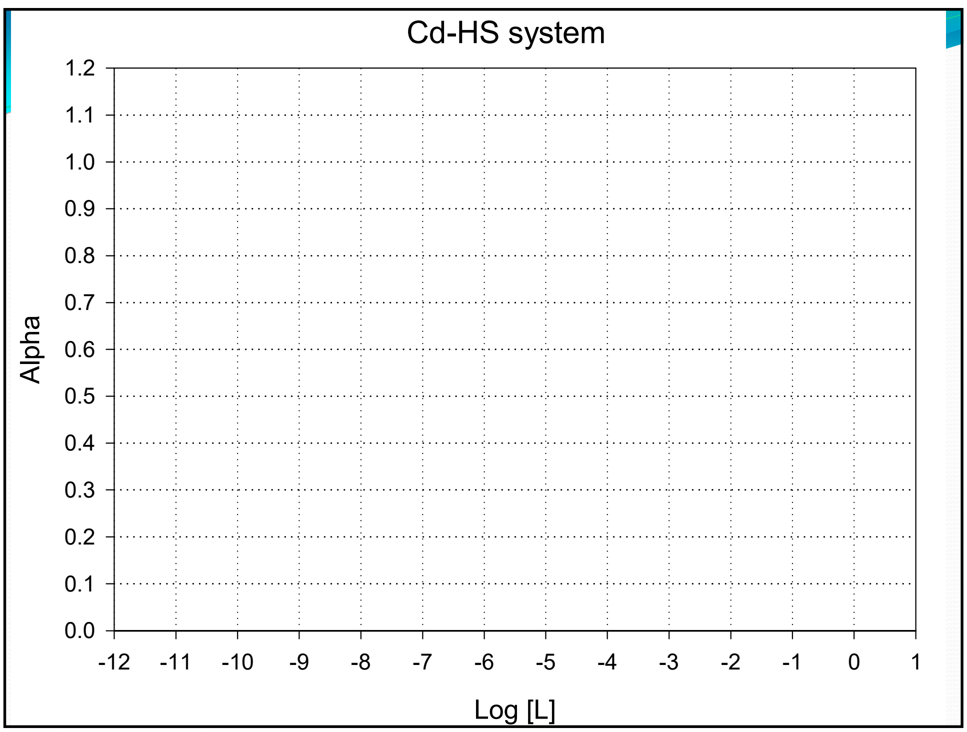
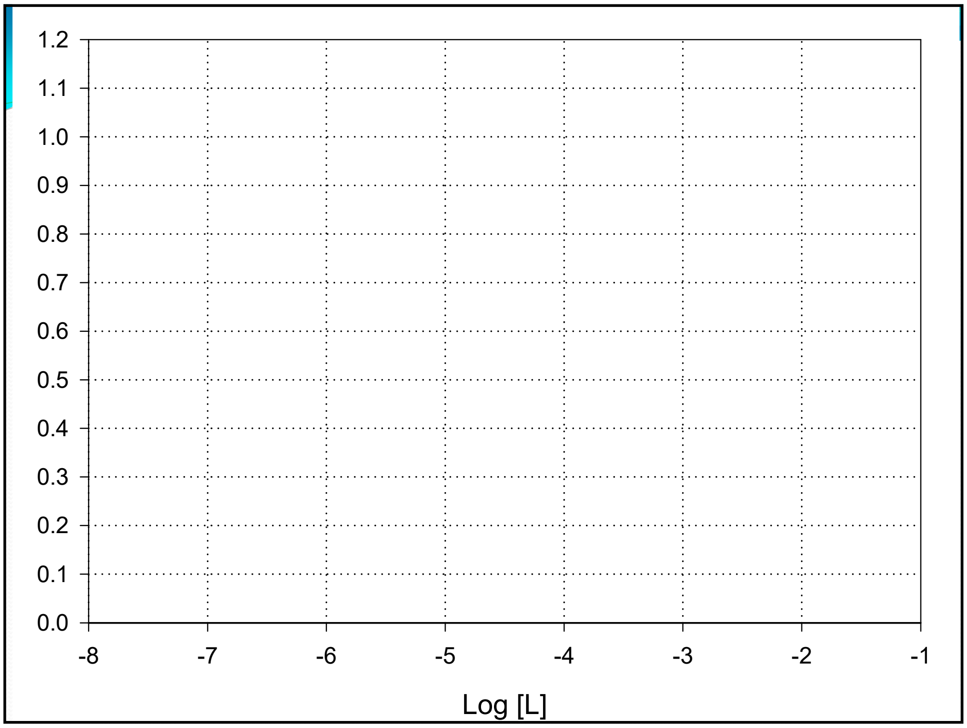
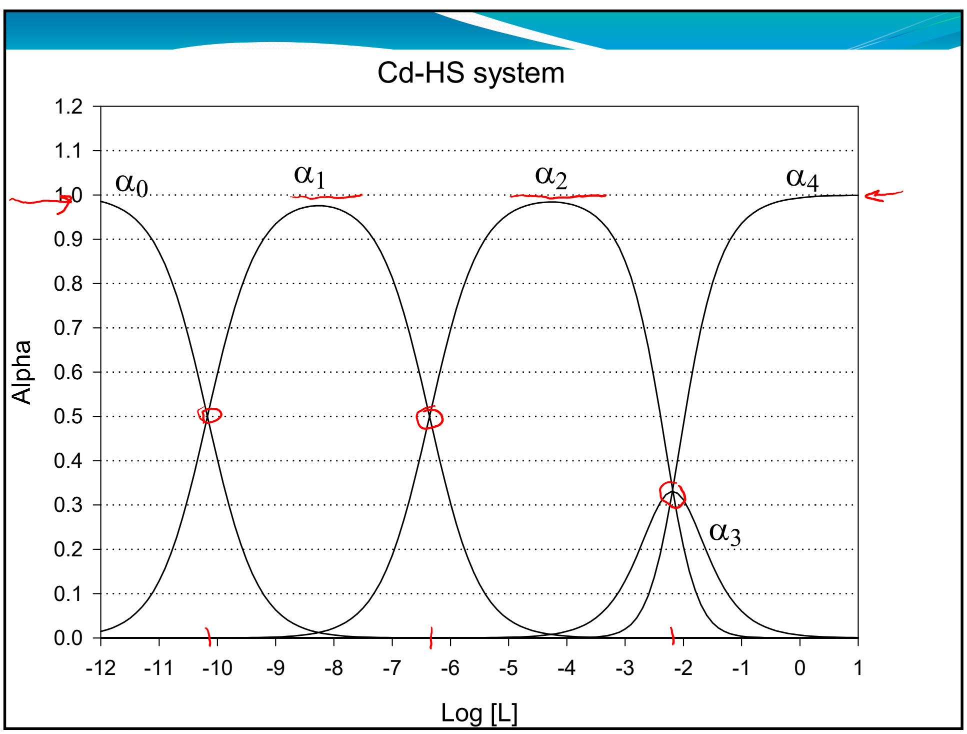
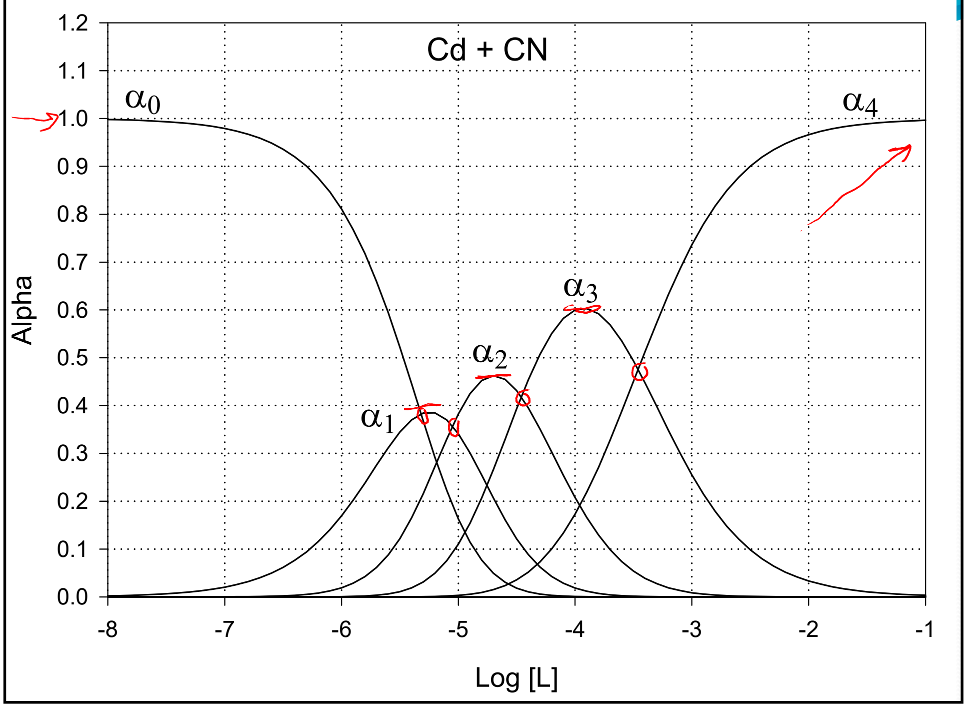



1.2
1.1
1.0
0.9
0.8
0.7
0.6
0.5
0.4
0.3
0.2
0.1
0.0
-8
-7
-6
-5
-4
-3
-2
-1
Log [L]

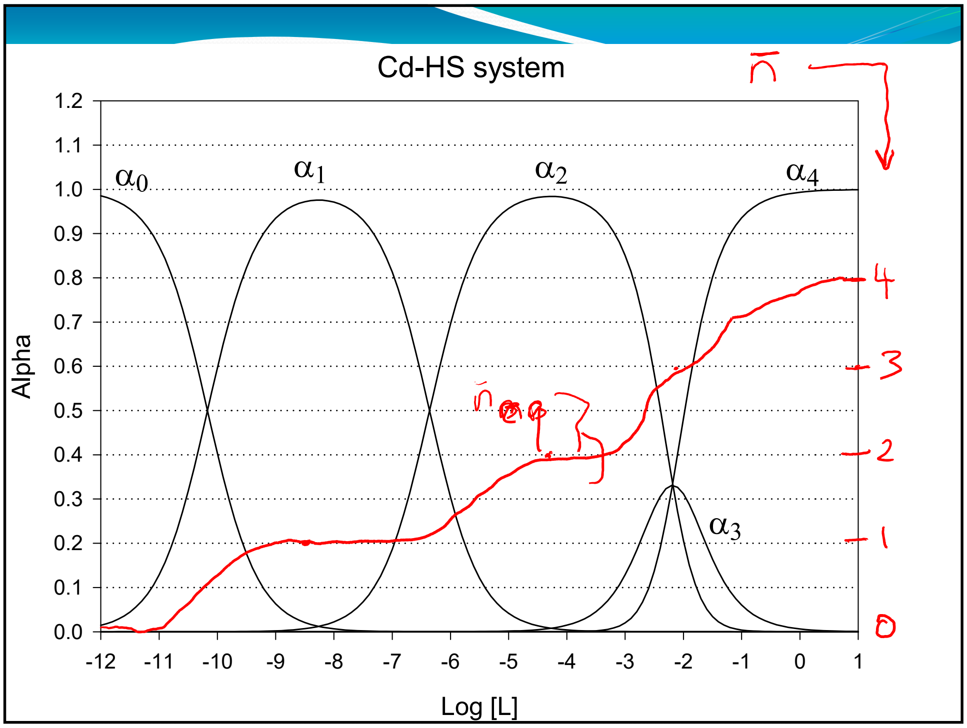

Cd-HS system
Alpha
Log [L]
α_0
α_1
α_2
α_3
α_4
1.2
1.1
1.0
0.9
0.8
0.7
0.6
0.5
0.4
0.3
0.2
0.1
0.0
-12
-11
-10
-9
-8
-7
-6
-5
-4
-3
-2
-1
0
1

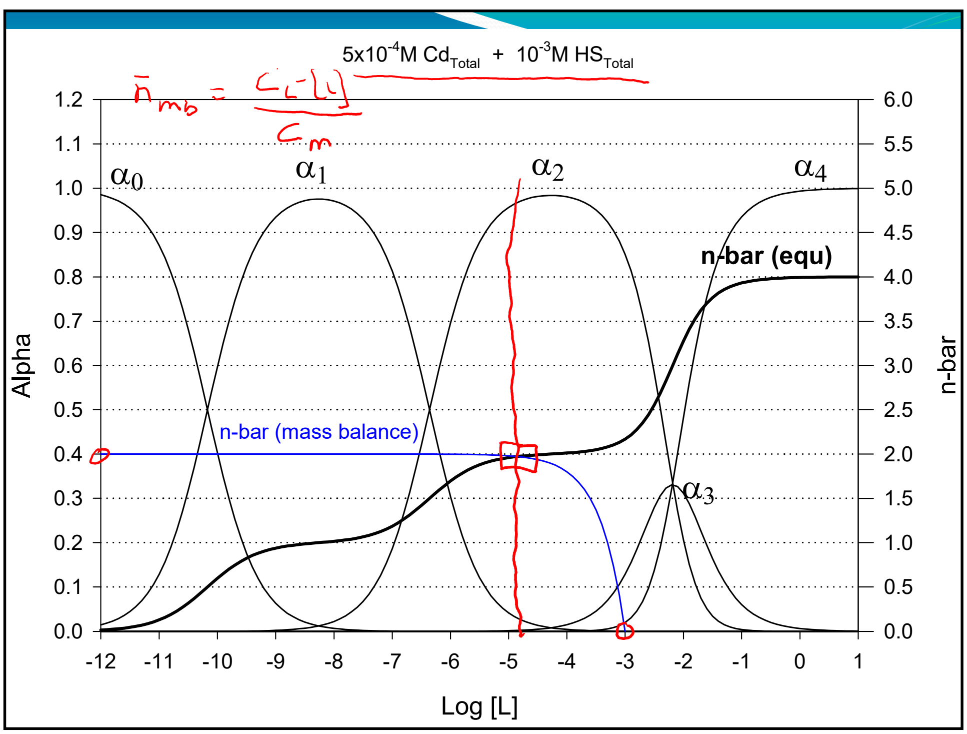


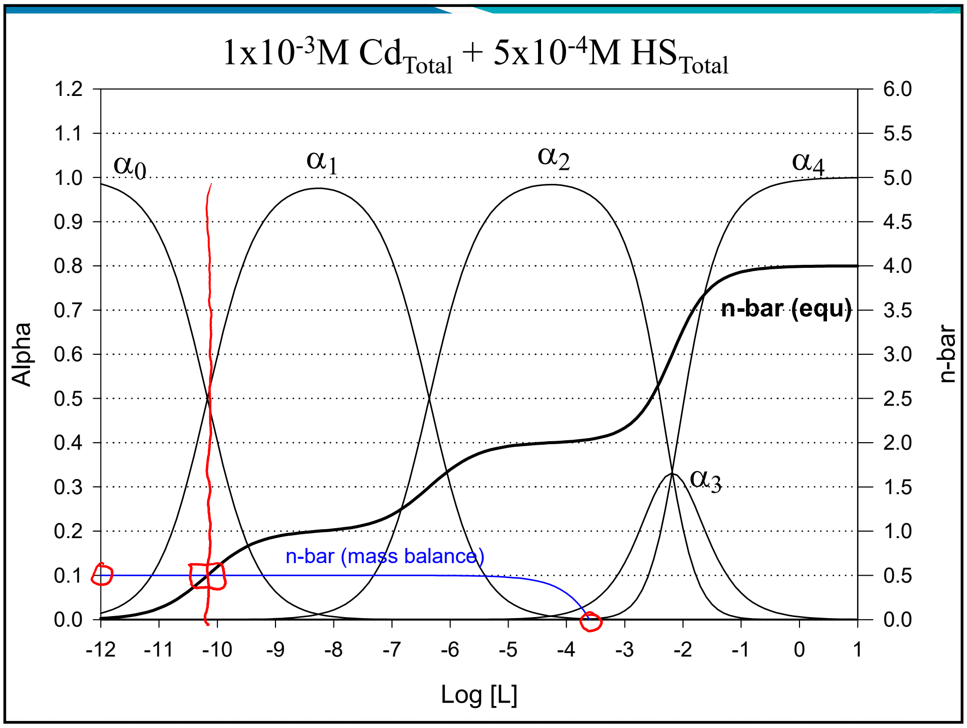
Cadmium Bisulfide

Species	log K	Log Beta
CdL	Log K_1 = 10.17	Log β_1 = 10.17
CdL_2	Log K_2 = 6.36	Log β_2 = 16.53
CdL_3	Log K_3 = 2.18	Log β_3 = 18.71
CdL_4	Log K_4 = 2.19	Log β_4 = 20.90

- Specific Problem
  - 5 x 10^{-4} M Cd total
  - 10^{-3} M HS total


Cd-HS system
Alpha
Log [L]
α0
α1
α2
α4
α3
n̄
n̄ eq
0
1
2
3
4

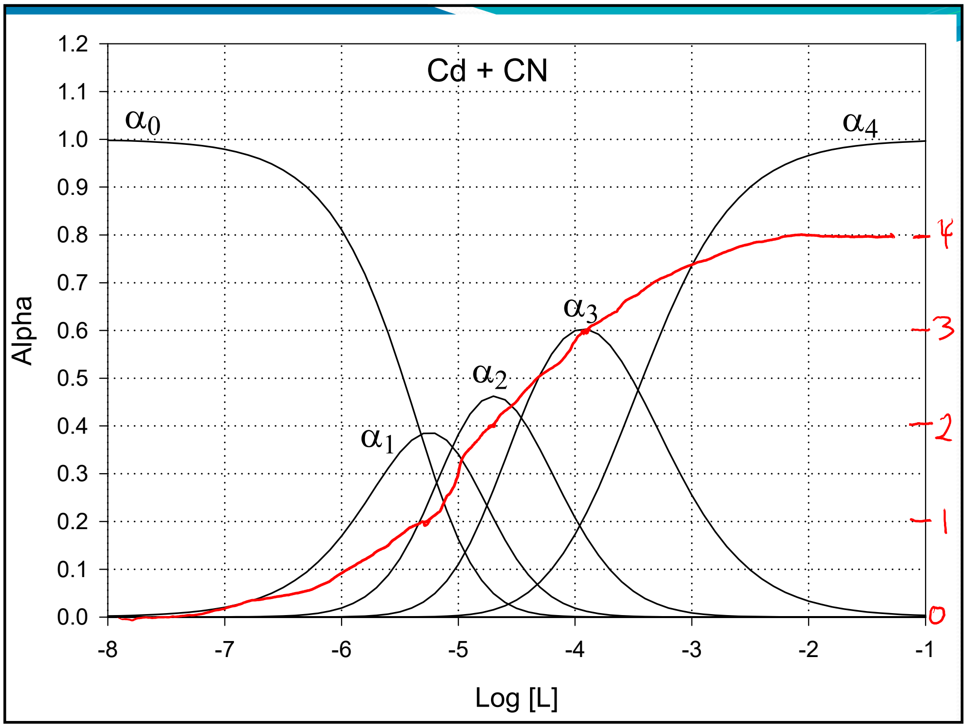

5x10-4M CdTotal + 10-3M HSTotal
n̄ mb = (CL - [L]) / Cm
Alpha
n-bar
Log [L]
α0
α1
α2
α4
α3
n-bar (equ)
n-bar (mass balance)

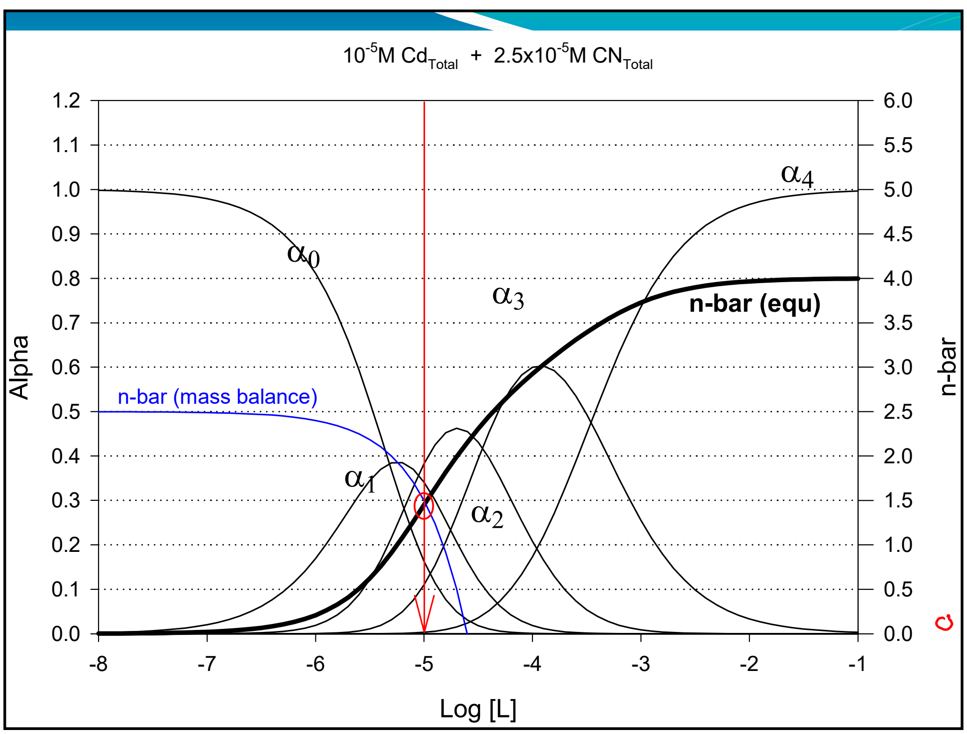


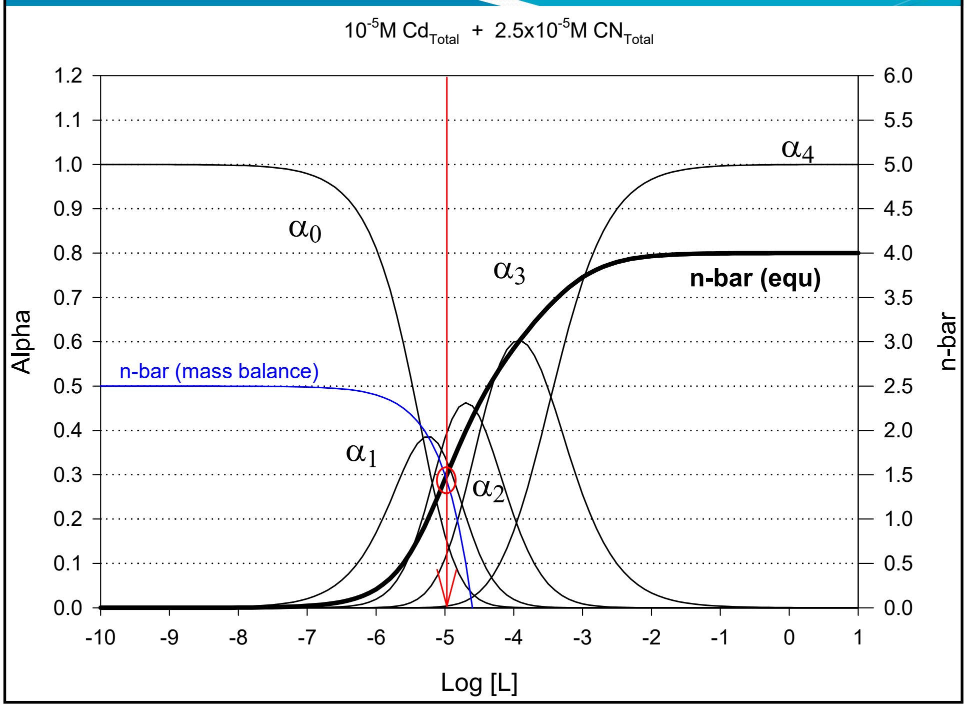
Cadmium Cyanide

Species	log K	Log Beta
CdL	Log K_1 = 5.32	Log β_1 = 5.32
CdL_2	Log K_2 = 5.05	Log β_2 = 10.37
CdL_3	Log K_3 = 4.46	Log β_3 = 14.83
CdL_4	Log K_4 = 3.46	Log β_4 = 18.29

- Specific Problem
  - 10^{-5} M Cd total
  - 2.5 x 10^{-5}M CN total


Cd + CN
Alpha
Log [L]
α0
α1
α2
α3
α4
4
3
2
1
0

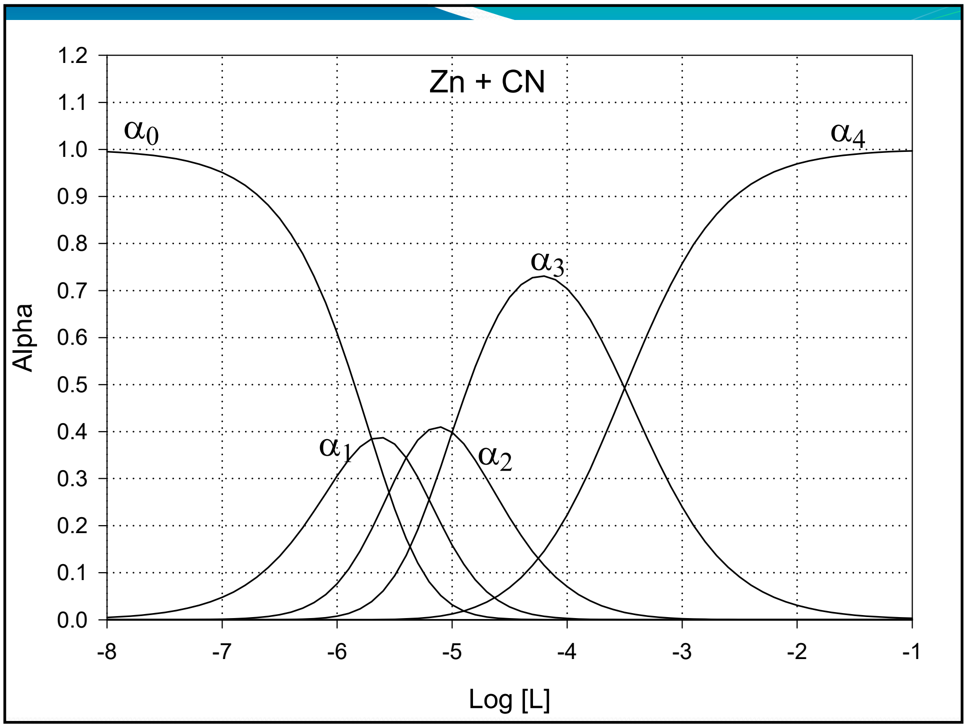

10-5M CdTotal + 2.5x10-5M CNTotal
Alpha
n-bar
Log [L]
α0
α1
α2
α3
α4
n-bar (equ)
n-bar (mass balance)

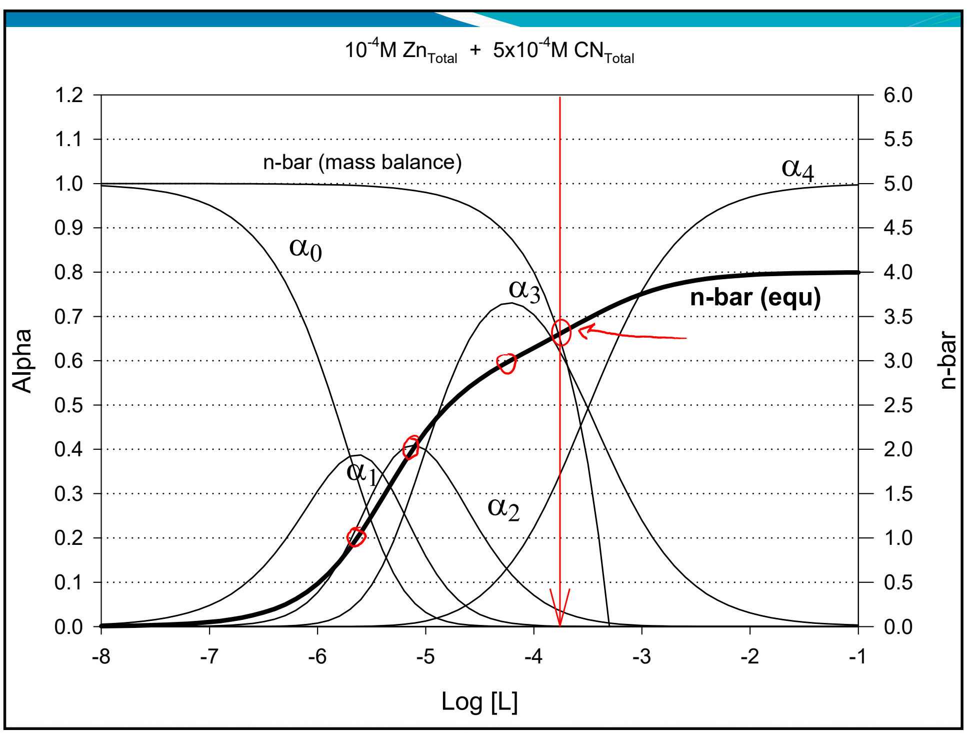


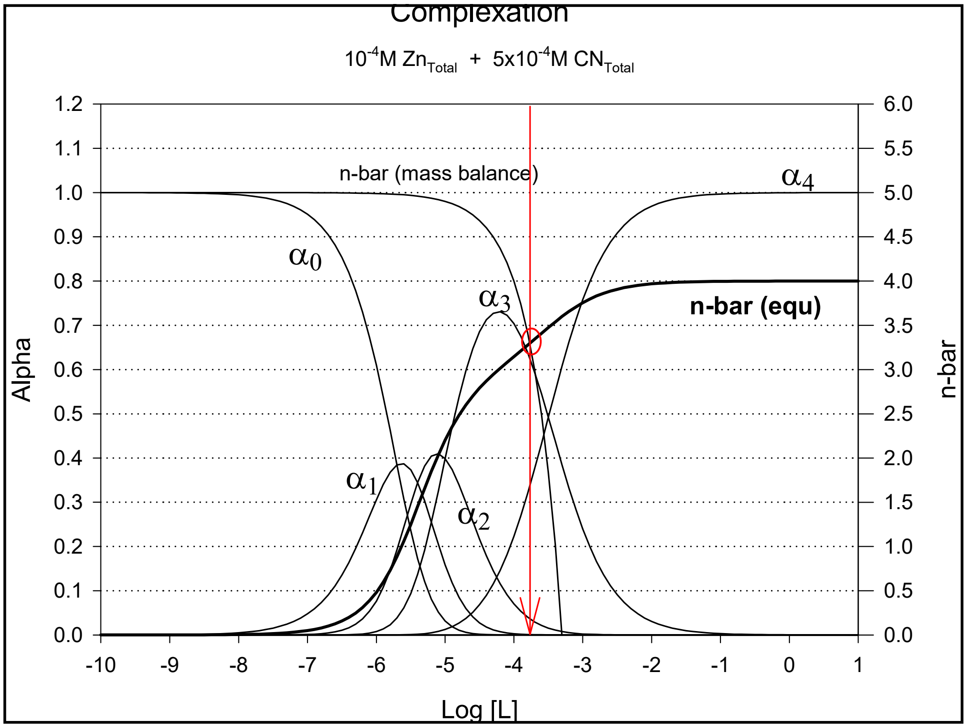
Zinc Cyanide

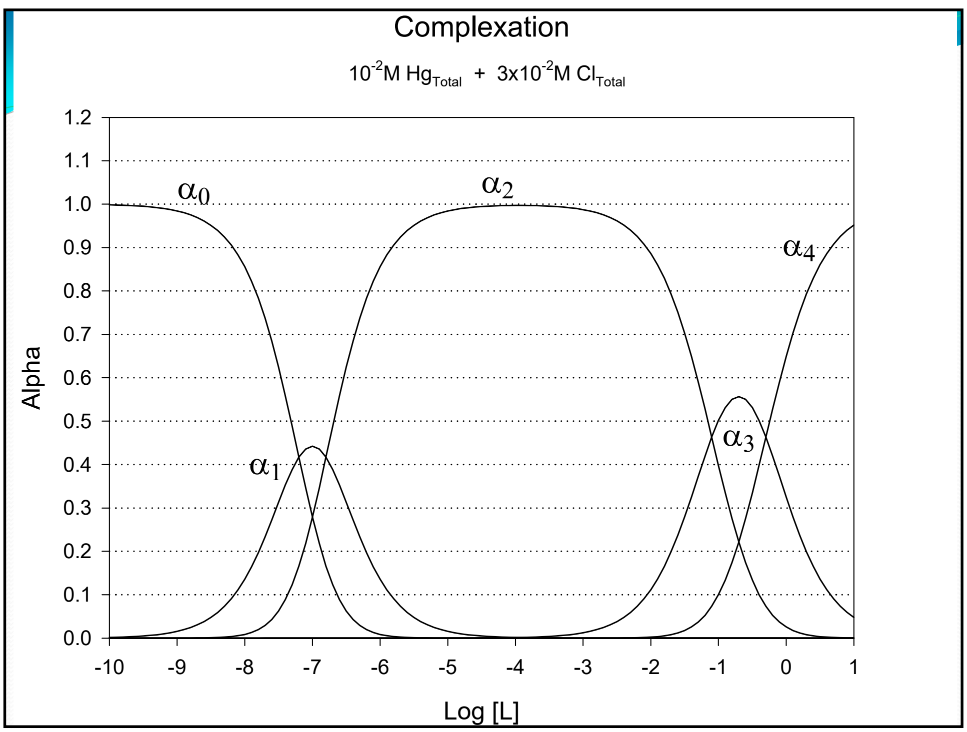
Species	Betas
ZnL	Log β_1 = 5.7
ZnL_2	Log β_2 = 11.1
ZnL_3	Log β_3 = 16.1
ZnL_4	Log β_4 = 19.6

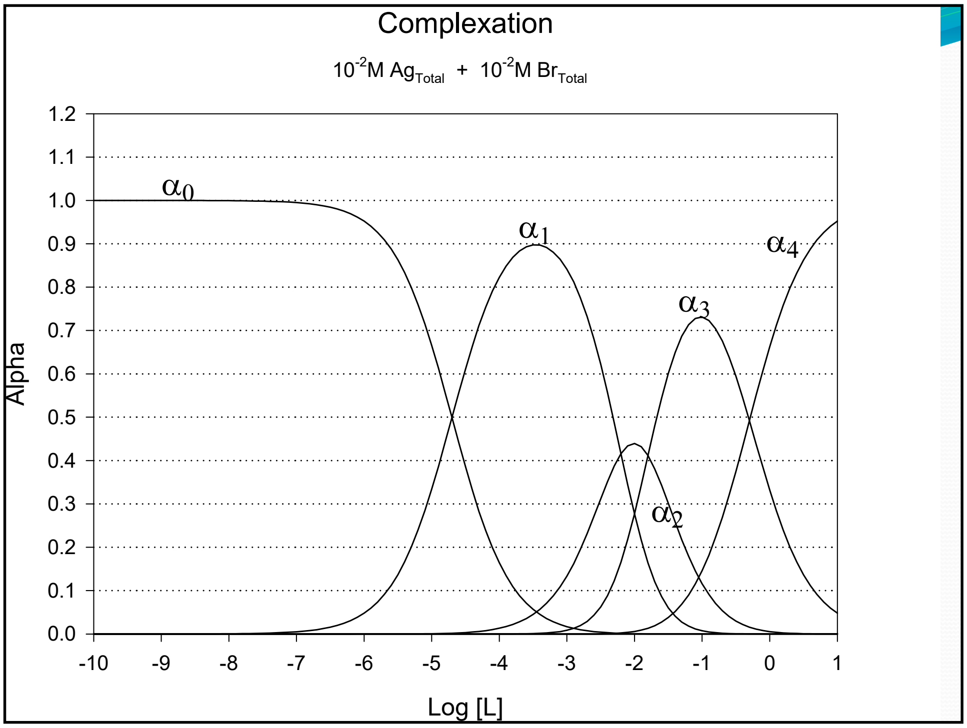
- Specific Problem
  - 10^{-4} M Zn total
  - 5 x 10^{-4}M CN total


Zn + CN
Alpha
Log [L]
α0
α1
α2
α3
α4


10-4M ZnTotal + 5x10-4M CNTotal
n-bar (mass balance)
n-bar (equ)
Alpha
n-bar
Log [L]
α0
α1
α2
α3
α4


Complexation
10^{-4}M Zn_{Total} + $5x10^{-4}$M CN_{Total}
n-bar (mass balance)
α_0
α_1
α_2
α_3
α_4
n-bar (equ)
Alpha
n-bar
Log [L]


Complexation
10^{-2}M Hg_{Total} + $3x10^{-2}$M Cl_{Total}
α_0
α_1
α_2
α_3
α_4
Alpha
Log [L]



Aluminum Fluoride system

- Significance
  - Aluminum in natural waters
  - Aluminum in coagulation
- Thermodynamic Values: Smith & Martel (Benjamin)
  - Log K_1 = 6.16 (7.01)
  - Log K_2 = 5.05 (5.74)
  - Log K_3 = 3.91 (4.27)
  - Log K_4 = 2.71 (2.70)
  - Log K_5 = 1.46 (1.08)
  - Log K_6 = 0 (-0.3)
- Now calculate alpha's

$$\alpha_0 \equiv \frac{[M]}{C_M} = \left(1 + \beta_1[L] + \beta_2[L]^2 + \cdots + \beta_n[L]^n\right)^{-1}$$

$$\alpha_n \equiv \frac{[ML_n]}{C_M} = \alpha_0 \beta_n [L]^n$$

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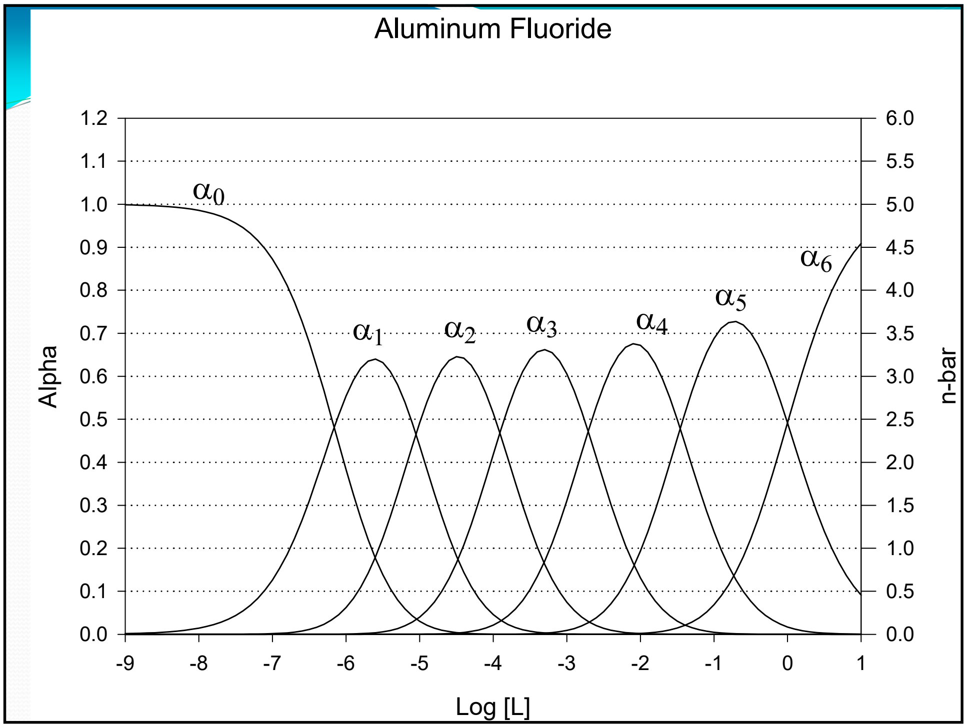
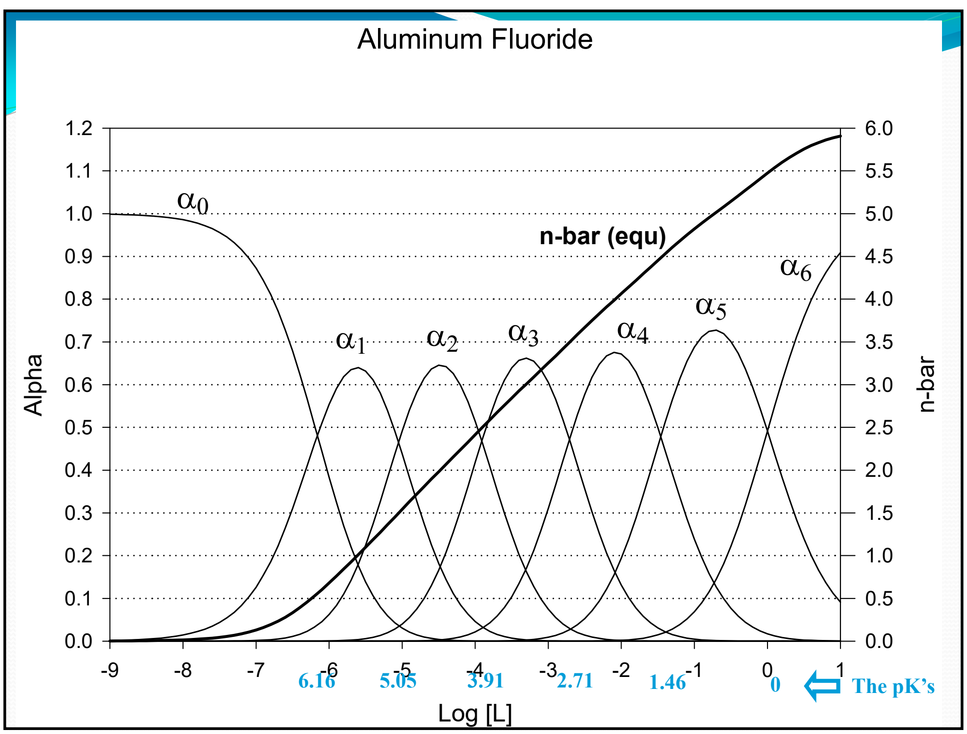
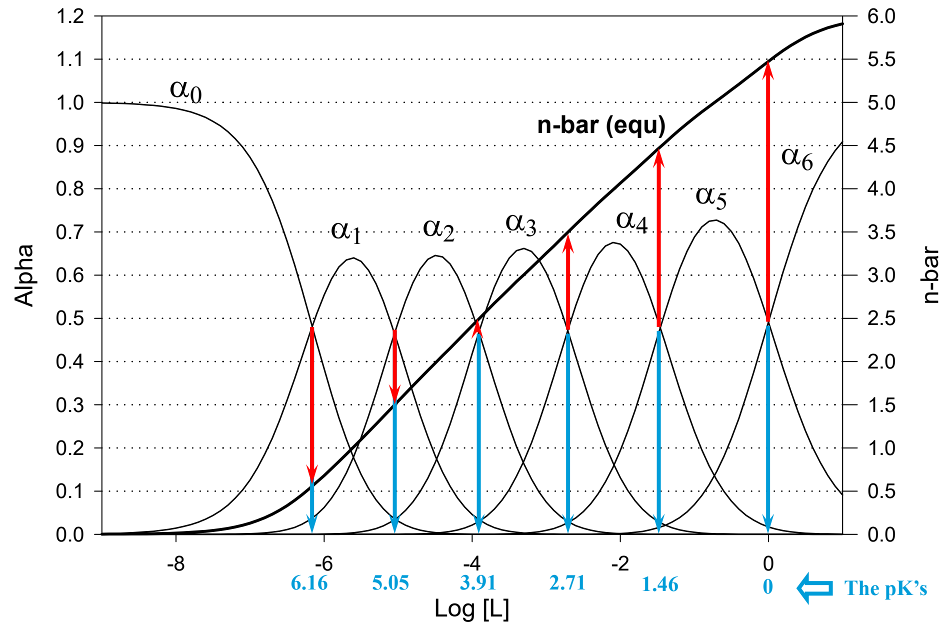

Aluminum Fluoride
Alpha
n-bar
Log [L]
α0
α1
α2
α3
α4
α5
α6


Aluminum Fluoride
Alpha
n-bar
Log [L]
n-bar (equ)
α0
α1
α2
α3
α4
α5
α6
6.16
5.05
3.91
2.71
1.46
0
The pK's



Natural Fluoride

Figure 1.4, pg. 9 in Benjamin, 2015

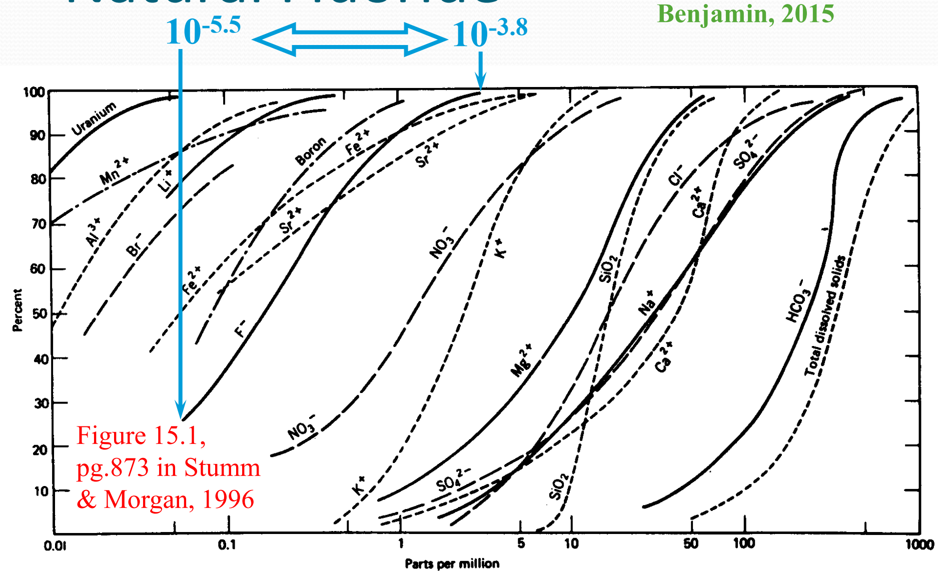
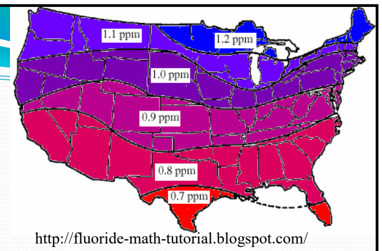


Figure 15.1, pg.873 in Stumm & Morgan, 1996

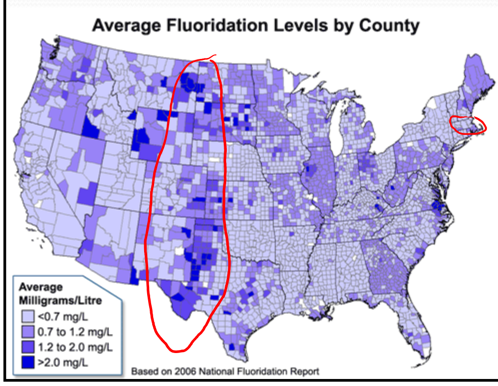
Fluoride addition

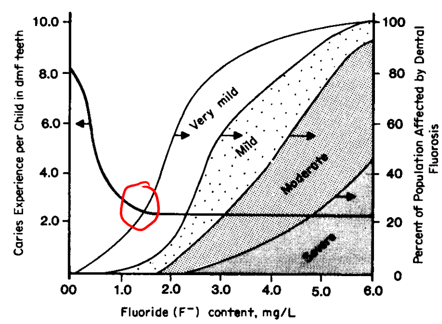
- Balance between Dental Caries and Fluorosis
- Recommended dose
  - 0.7 to 1.2 mg/L, Based on temperature



http://fluoride-math-tutorial.blogspot.com/

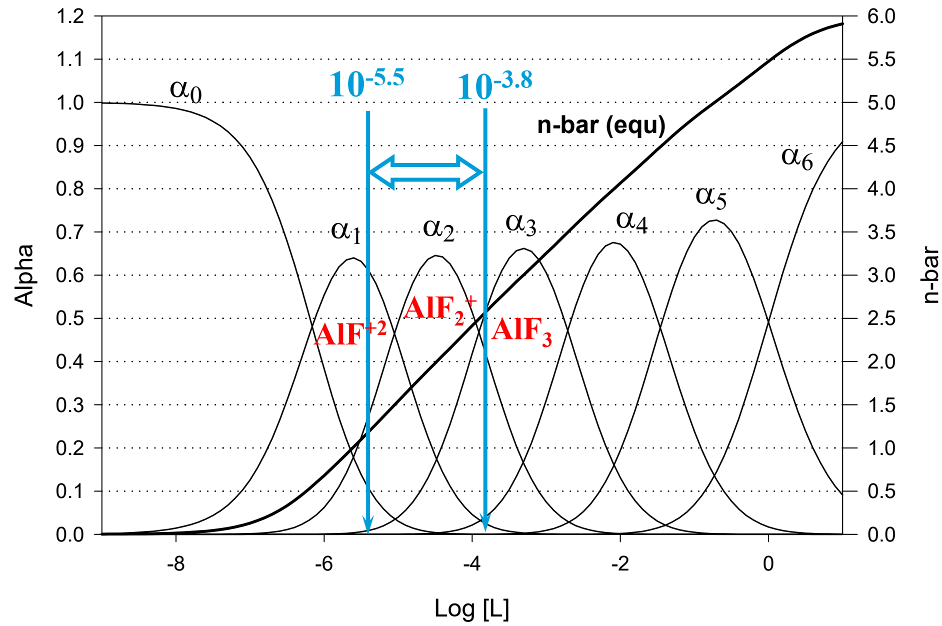
Fig. 15.3 from Water Quality & Treatment, 1999 (5th edition)





Al-F Problems & Discussion

- Typical WT Situation
  - Alum dose = 33 mg/L
  - Total Fluoride = 1.9 mg/L
- High Fluoride pulse & high alum dose
  - Alum dose = 660 mg/L
  - Total Fluoride = 190 mg/L
- Impacts of OH complexation?



Iron Thiocyanate system

- Significance
  - Metal plating wastewaters
  - Used in colorimetric analysis of iron
- Thermodynamic Values
  - Log K_1 = 2.11
  - Log K_2 = 1.19
  - Log K_3 = 0
  - Log K_4 = 0
  - Log K_5 = -0.1
  - Log K_6 = -0.9
- Now calculate alpha's

$$\alpha_0 \equiv \frac{[M]}{C_M} = \left(1 + \beta_1[L] + \beta_2[L]^2 + \cdots + \beta_n[L]^n\right)^{-1}$$

$$\alpha_n \equiv \frac{[ML_n]}{C_M} = \alpha_0 \beta_n [L]^n$$

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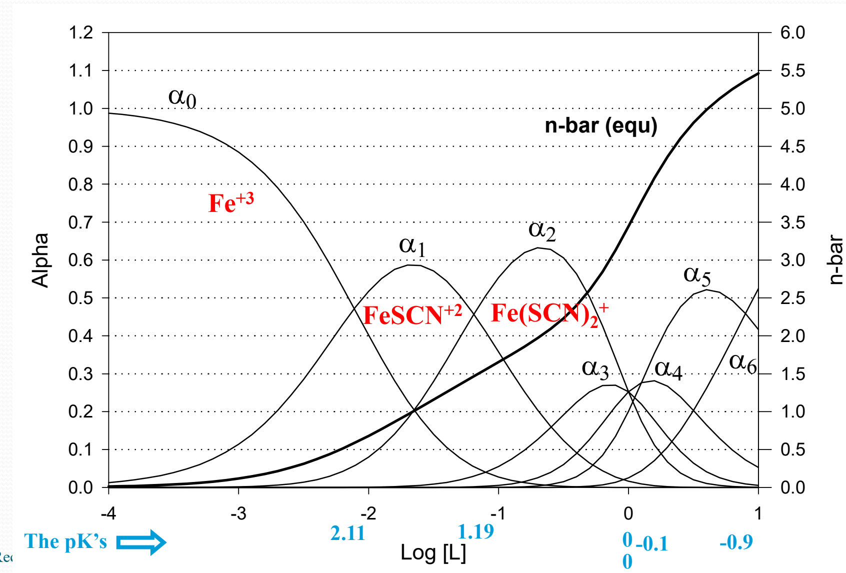


Specific problem

- Total concentrations
  - C_M = 0.1 M
  - C_L = 0.1 M
- Mass based equation
  - N-bar = 1-10[SCN^-]
- Solution: n-bar = 0.85
  - [Fe^{+3}]=0.028M
  - [$FeSCN^{+2}$] = 0.057M
  - [$Fe(SCN)_2^+$] = 0.014M

Fe-S problem

- Below are the equilibria for the Fe^{+2} – HS system as listed in Benjamin's book. Note that are no equilibria for FeL, as this species is never significant. Prepare a graph of alpha values (vs log[HS-]) for the this system. Using this graph determine the complete ferrous-iron speciation in groundwater where the total sulfide concentration is 0.2 mM and total ferrous iron is 0.1 mM. Assume the pH of the groundwater is about 8.

Species	Ligand	
	HS-	PO_4^{-3}
FeL_2	8.95	
FeL_3	10.99	
FeH_2L		22.25

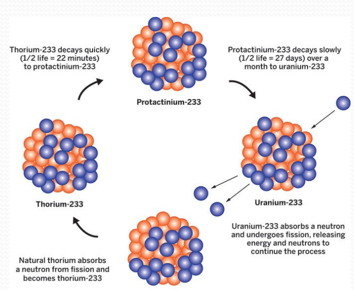
Log ß values (From Table 8.3 ; pg 374)

Thorium

- ^{232}Th is 99.98% of natural abundance
  - weakly radioactive ($t_{1/2}$ = 14 billion yrs)
  - Most abundant radioactive element in nature
- Uses:
  - Nuclear power: forms ^{233}U
    - $^{232}_{90}Th + n \rightarrow ^{233}_{90}Th + \gamma \rightarrow ^{233}_{91}Pa \rightarrow ^{233}_{92}U$
- Health effects:
  - Bone, liver and lung cancer
- Solvated by 9 waters (like many Lanthanide and Actinide elements

No specific standard, but EPA has established a Maximum Contaminant Level (MCL) of 15 picoCuries per liter (pCi/L) for alpha particle activity, excluding radon and uranium, in drinking water.

1 Ci = $3.7x10^{10}$ dps

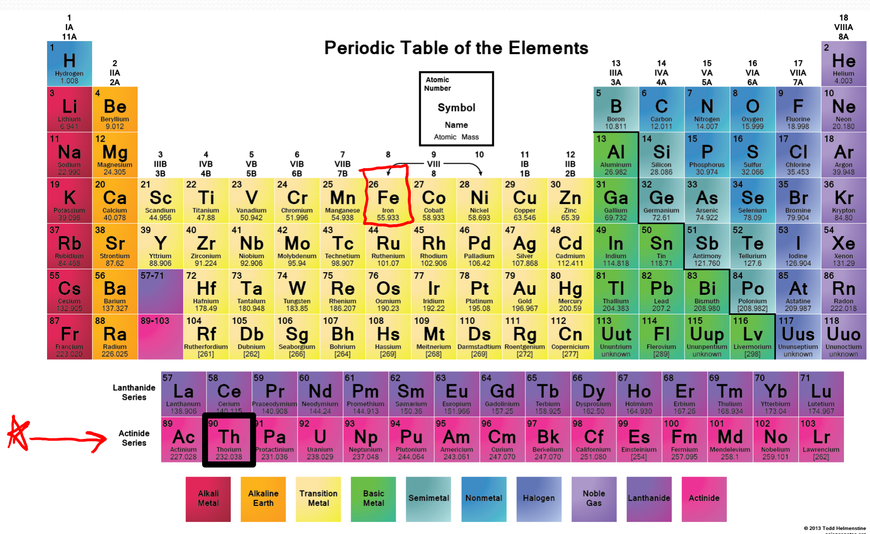


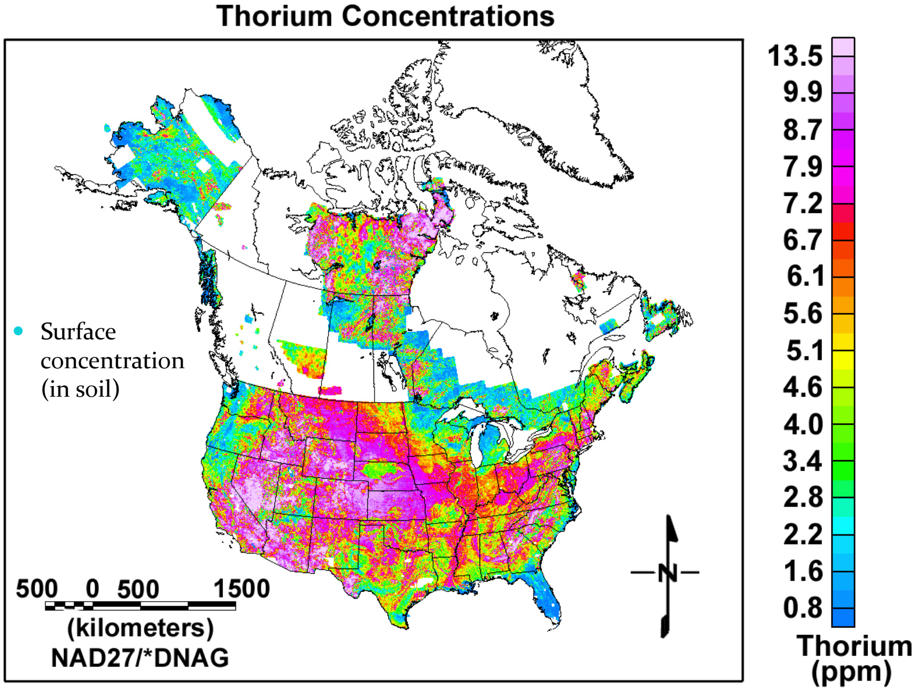
Converting pCi/L to ug/L

- Spreadsheet calculation for Thorium-232

Thorium	232	AMU										
		EPA MCL	15	pCi/L =	1.5E-11	Ci/L		1	Ci =	3.7E+10	d/s	
					0.555	d/s/L						
		Th 232										
			t 1/2 =	14000000000	years							
			k =	4.95105E-11	per year				Avogadro's #	6.02E+23	atoms/mole	
			disintegration rate=	2.98053E+13	d/mole/yr =	944473.9	d/mole/s					
			conc =	5.87629E-07	moles/L =	0.587629	uM =	136	ug/L			

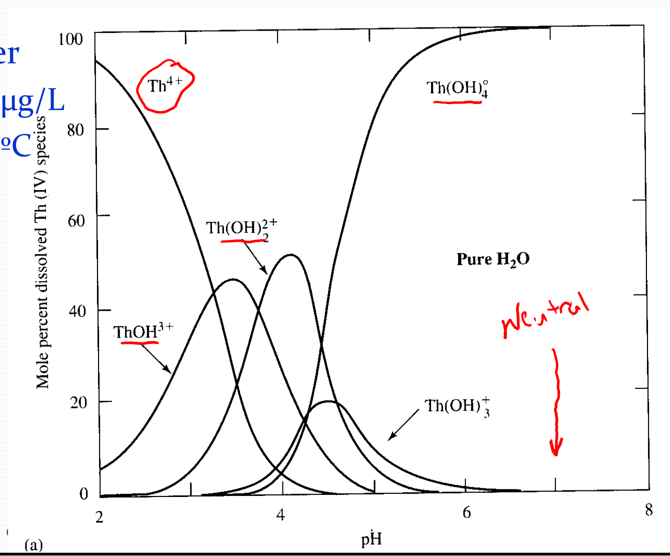
Thorium: Actinide Element





Thorium example I

- In pure water
  - C_{Th} = 0.01 μg/L
  - Temp = 25ºC



From: Langmuir, 1997; Fig. 3.12a,
Original source: Langmuir & Herman, 1980; *Geochim. Et Cosmochim. Acta* 44(11)1753-1766

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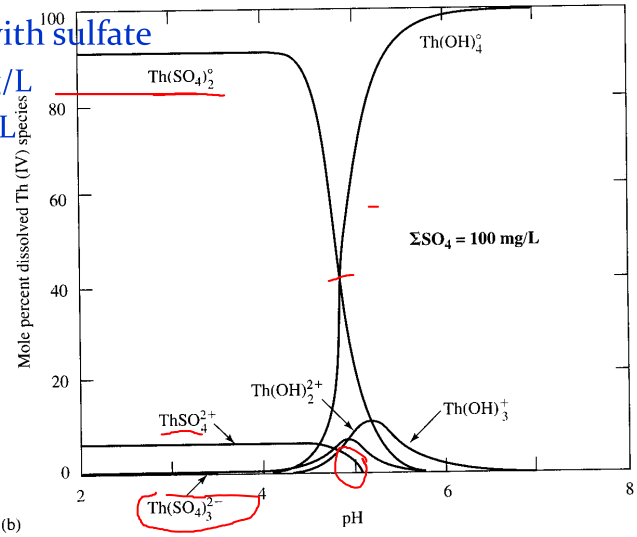
Thorium example II

- In pure water with sulfate
  - C_{SO4} = 100 mg/L
  - C_{Th} = 0.01 µg/L
  - Temp = 25ºC



From: Langmuir, 1997, Fig. 3.12b

Original source: Langmuir & Herman, 1980; *Geochim. Et Cosmochim. Acta* 44(11)1753-1766

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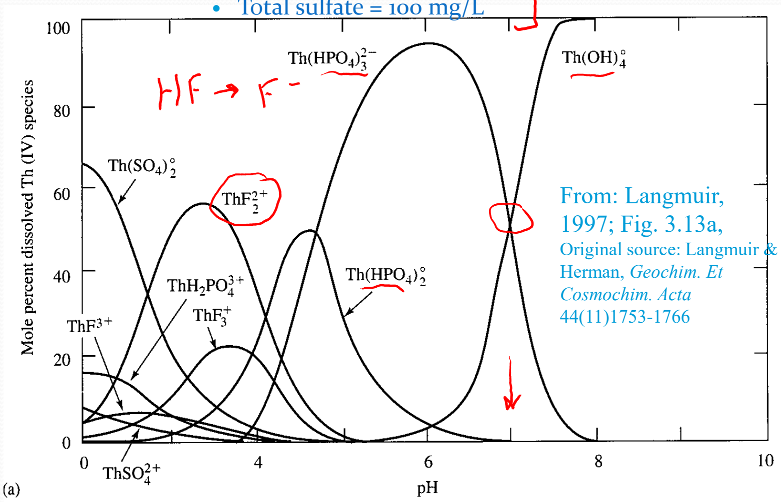
Thorium example III

- Model groundwater without organics
  - C_{Th} = 0.01 µg/L & Temp = 25ºC
  - Groundwater composition
    - Total fluoride = 0.3 mg/L
    - Total chloride = 10 mg/L
    - Total phosphate = 0.1 mg/L
    - Total sulfate = 100 mg/L

Up to 220 mg-Th/kg-P in phosphate rock, less in P from sewage sludge

Th-P solids are less bioavailable



From: Langmuir, 1997; Fig. 3.13a, Original source: Langmuir & Herman, *Geochim. Et Cosmochim. Acta* 44(11)1753-1766

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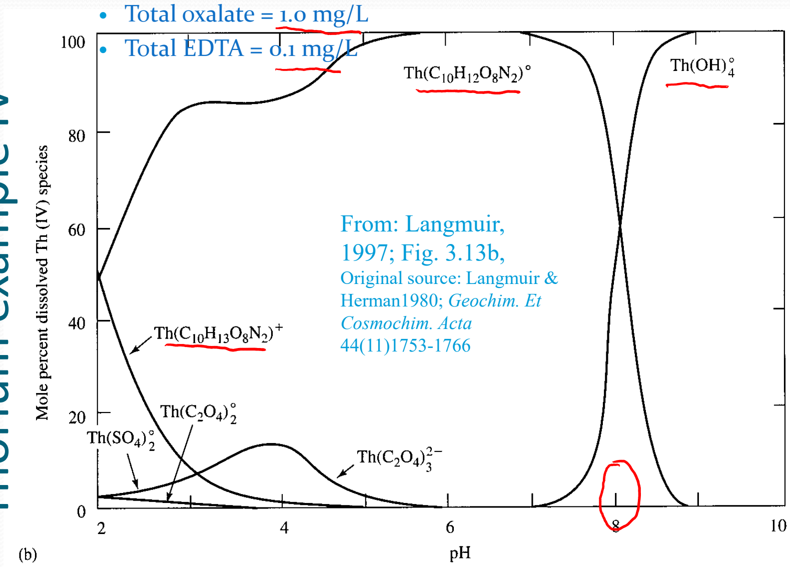
Thorium example IV

- Model groundwater with organics
  - Same inorganic groundwater composition
  - With organics
    - Total oxalate = 1.0 mg/L
    - Total EDTA = 0.1 mg/L



From: Langmuir, 1997; Fig. 3.13b,
Original source: Langmuir & Herman1980; *Geochim. Et Cosmochim. Acta* 44(11)1753-1766

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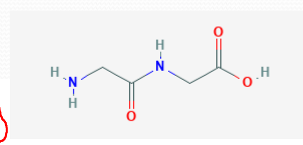
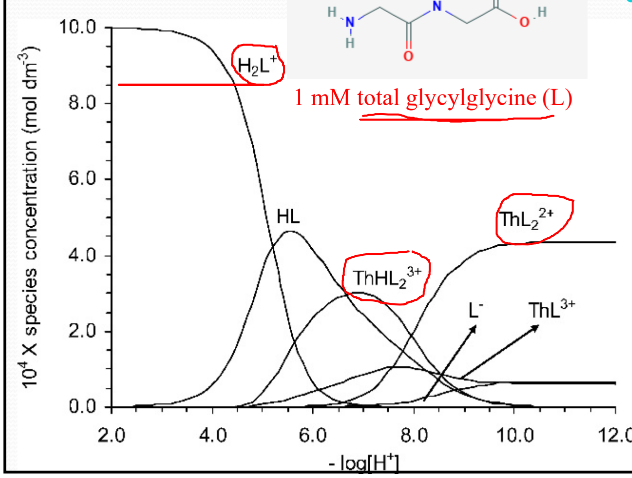
In the human body



1 mM total glycylglycine (L)



- "The thorium(IV) ion readily reacts with in vivo amino acids, peptides, nucleic acids, proteins, etc. to form stable complexes which are distributed in the body, primarily in the liver, bone, and kidneys"

Kiani et al., 2009 J. Chem. Eng. Data, 54(12) 3247-3251, "Complex Formation of Thorium(IV) Ion with Glycyl-Glycine and Glycyl-Valine"

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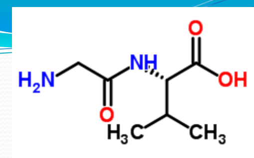
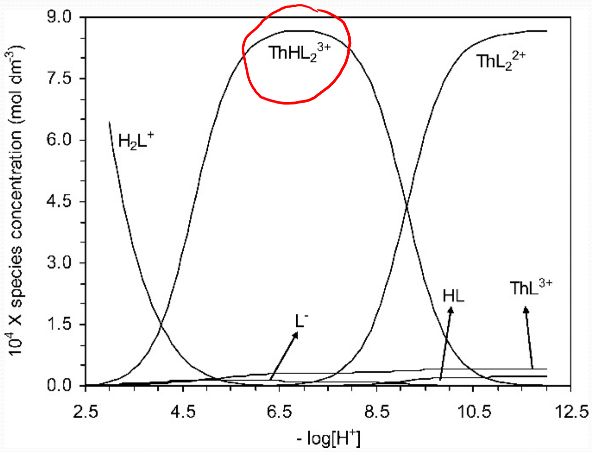
Another dipeptide



- This time with 1 mM total glycylvaline (L)



Kiani et al., 2009 J. Chem. Eng. Data, 54(12) 3247-3251, "Complex Formation of Thorium(IV) Ion with Glycyl-Glycine and Glycyl-Valine"

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End

- To next lecture

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Fe-S problem

- Below are the equilibria for the Fe^{+2} – HS system as listed in Benjamin's book. Note that are no equilibria for FeL, as this species is never significant. Prepare a graph of alpha values (vs log[HS-]) for the this system. Using this graph determine the complete ferrous-iron speciation in groundwater where the total sulfide concentration is 0.2 mM and total ferrous iron is 0.1 mM. Assume the pH of the groundwater is about 8.

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FeL_2	8.95	
FeL_3	10.99	
FeH_2L		22.25

Log ß values (From Table 8.3 ; pg 374)

Calculations

Log ß values (From Table 8.3 ; pg 374)

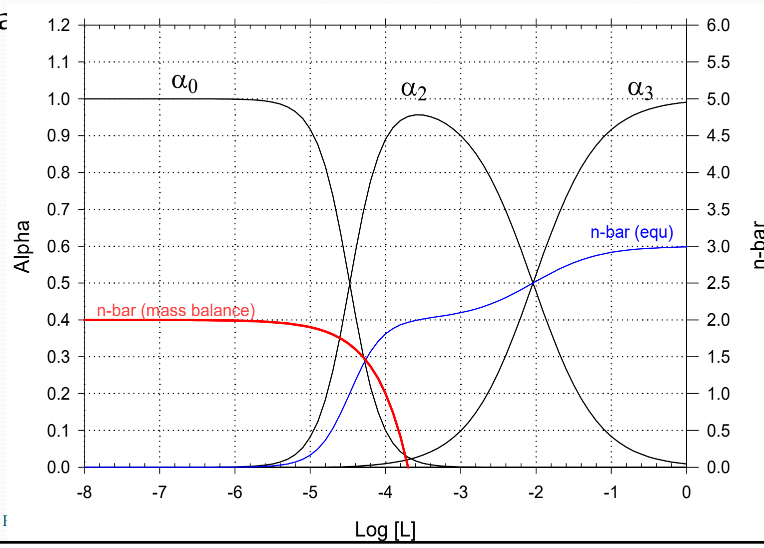
Species	Ligand	
	HS-	PO_4^{-3}
FeL_2	8.95	
FeL_3	10.99	
FeH_2L		22.25

- calculations

$$\beta_2 = 10^{8.95} = \frac{[FeL_2]}{[Fe]\ [L^2]}$$

$$\beta_3 = 10^{10.99} = \frac{[FeL_3]}{[Fe]\ [L^3]}$$

- a



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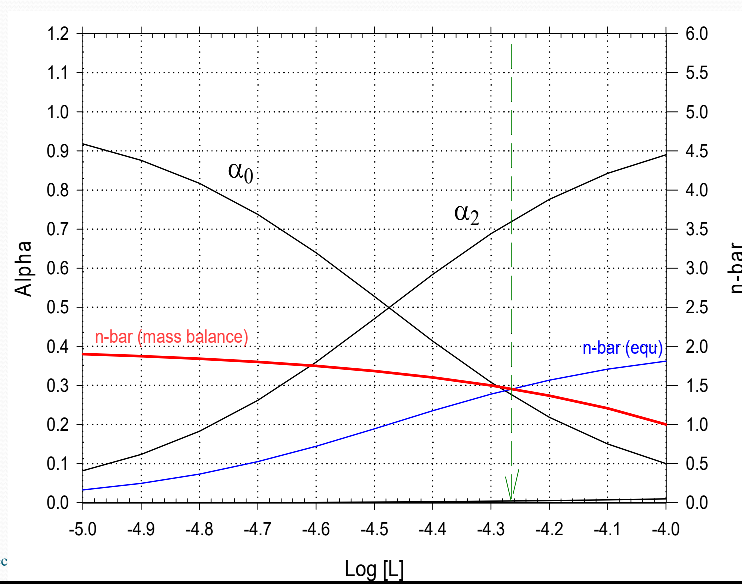

David Rec

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Species	Conc (M)	Log C
HS-	5.37E-05	-4.27
Fe^{+2}	2.70E-05	-4.57
$Fe(HS)_2^o$	7.30E-05	-4.14
$Fe(HS)_3^-$	0	
H^+	1.00E-08	-8
OH^-	1.00E-06	-6
H_2S	5.75E-06	-5.24
S^{-2}	5.37E-11	-10.27